



Solution of Questions For Short Answer

Chapter 25 : Calorimetry

Ans.1.

Heat is not conserved quantity as heat only exists when there is some energy transferred.

Ans. 2.

As cal is very small unit energy may be calculated in Kcal so in this case difference between 4.186 and 4 becomes large.

Ans. 3.

It is necessary because heat from environment may absorbed by the calorimeter or heat of the object may be released by the calorimeter in environment so as to measure precise quantity of heat.

Ans. 4.

Heat is measure of transfer of energy there is no heat in the body or heat of the body. so heat of the two objects doesn't make any sense.

Ans. 5.

Inlet at the top allows material to got down due to gravity and material stays more time in steam chamber.

Ans. 6.

The energy is consumed in the process of changing state of the material so as from liquid to vapor or solid to liquid. the heat involved changes form of material and the temperature doesn't increases.

Ans. 7.

Specific heat capacity of melting ice= 4.210 (as ice starts melting above 0°c)
specific heat capacity of boiling water=4.219 (as water starts boiling at 100°c)

Ans. 8.

The skin felts more burning as latent heat is involved in steam at 100°c and thus total heat of steam is greater than water at 100°c



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Ans. 9.

As the specific heat of water is higher than soil or sand thus lesser heat is absorbed or released than stones or sand. Thus atmospheric temperature changes in very less amount.

Ans.10.

Thermometer bulb must have small heat capacity so as with very small amount of heat the temperature rises up to prominent amount.

